

CLASSIFICATION OF ELEMENTS, CHEMICAL BONDING, PERIODICITY IN PROPERTIES & PHYSICAL TRANSFORMATION BASED GENERAL SCIENCE MCQ PRACTICE QUESTIONS AND ANSWERS PDF WITH EXPLANATION

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Q1. Consider the following statements regarding **Bohr atomic model**:

- It introduces the idea of stationary orbits.
- It assumes that the angular momentum of the electron is equal to $-(1/2)h/2\pi$
- It uses the planetary model of the atom involving circular orbits.

Which of the statements given above are **correct**?

- a) 1 and 3
 - b) 1 and 2
 - c) 2 and 3
 - d) 1, 2 and 3
-

Q2. Which of the following are values of gas constant R ?

- $0.0821 \text{ L atm k}^{-1}\text{mol}^{-1}$
- $8.21 \text{ L torr k}^{-1}\text{mol}^{-1}$
- $82.1 \text{ atm mL k}^{-1}\text{mol}^{-1}$
- $8.314 \text{ J mol}^{-1}\text{k}^{-1}$

- a) 1, 2 and 4
 - b) 1, 2 and 3
 - c) 1, 3 and 4
 - d) 2, 3 and 4
-

Q3. Element A belongs to Group VII in p-block and element B belongs to Group I in s-block of the periodic table. Out of the following assumptions, the correct one is :

- a) A and B are non-metals

- b) A is a metal and B is a non-metal
c) A and B are metals
d) A is a non-metal and B is a metal
-

Q4. Why does water boil at a temperature below 100°C at higher altitudes?

- a) The gravitational attraction is less
b) Because of heavy winds in mountains
c) The atmospheric pressure decreases and hence boiling point is lowered
d) None of the above is correct

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Q5. Which of the following is/are the example/examples of chemical change?

- Crystallization of sodium chloride
- Melting of ice
- Souring of milk

Select the correct answer using the code given below.

- a) Only 3
b) 1, 2 and 3
c) Only 1 and 2
d) None of the above
-

Q6. The early attempt to classify elements as metals and nonmetals was made by

- a) Lothar Meyer
b) Lavoisier
c) Mendeleev
d) Henry Moseley
-

Q7. Example of corrosion is

- a) Green coating on Copper
 - b) Brown coating on Iron
 - c) Black coating on Silver
 - d) Above three
-

Q8. Which among the following atoms has high atomic radi?

- a) Rb
 - b) K
 - c) Na
 - d) Cs
-

Q9. The scientist who made maximum contribution towards periodic table was

- a) Rutherford
 - b) Dalton
 - c) Chadwick
 - d) Mendeleev
-

Q10. According to Newland's law of octaves, which element is the repetition of the first element in the periodic table?

- a) Nitrogen
 - b) Chlorine
 - c) Oxygen
 - d) Sulphur
-

Q11. To which block is related an element having electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^1$ in the periodic table—

- a) p-block
 - b) d-block
 - c) s-block
 - d) f-block
-

Q12. Consider the following statements with reference to the periodic table of chemical elements:

1. Ionisation potential gradually decreases along a period
2. In a group of elements, electron affinity decreases as the atomic weight increases
3. In a given period, electronegativity decreases as the atomic number increases

Which of these statement (s) is/are correct?

- a) 2 only
- b) 1 and 3
- c) 1 only
- d) 2 and 3

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Q13. The noble gases are unreactive because

- a) they have a full outer shell of electrons
 - b) they have a half outer shell of neutrons
 - c) they react with sodium
 - d) they are too thin
-

Q14. Which of the following elements does not lose an electron easily?

- a) F
- b) Mg

c) Na

d) Al

Q15. Which of the following are correct statements?

- The canonical structures have no real existence
- Every AB_5 molecule does in fact have square pyramidal structure
- Multiple bonds are always shorter than corresponding single bonds
- The electron-deficient molecules can act as Lewis acids

a) 2, 3 and 4

b) 1, 2 and 3

c) 1, 3 and 4

d) 1, 2 and 4

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Answers to the above questions :

Q1. Answer: (a)

According to Bohr, model electrons revolve around the nucleus in stationary orbits.

The angular momentum of these orbits is an integral multiple of $\frac{h}{2\pi}$.

It uses a planetary model of the atom involving circular orbits.

Q2. Answer: (c)

Q3. Answer: (d)

Element A belongs to the halogens (Group VII) group and is a non-metal.

While element B belongs to the alkali metal group (Group I) and is a metal.

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Q4. Answer: (c)

Q5. Answer: (a)

Q6. Answer: (b)

Q7. Answer: (d)

Q8. Answer: (d)

Q9. Answer: (d)

Q10. Answer: (c)

Q11. Answer: (c)

Q12. Answer: (a)

Q13. Answer: (a)

Q14. Answer: (a)

F has a tendency to gain an electron.

Q15. Answer: (c)

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