

# CLASSIFICATION OF ELEMENTS, CHEMICAL BONDING, PERIODICITY IN PROPERTIES & PHYSICAL TRANSFORMATION BASED GENERAL SCIENCE MCQ PRACTICE QUESTIONS AND ANSWERS PDF WITH EXPLANATION

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Created By [Careericons](#) Team

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**Q1.** The elements with atomic numbers 2, 10, 18, 36, 54 and 86 are all

- a) noble gases
  - b) noble metals
  - c) halogen
  - d) light metals
- 

**Q2.** Change of water into the vapour is called

- a) Physical
  - b) Chemical
  - c) Natural
  - d) Biological
- 

**Q3.** Which of the following statements is incorrect from the point of view of modern periodic table ?

- a) There are eighteen vertical columns called groups
  - b) Transition elements fit in the middle of long periods
  - c) Elements are arranged in the order of increasing atomic number
  - d) Noble gases are arbitrarily placed in eighteenth group
- 

**Q4.** Which fact is not valid for Dobereiner's triads?

- a) The properties of middle element is roughly average of the other two elements

- b) The elements of triads belong to the same group of modern periodic table
- c) The atomic weight of middle element is roughly average of the other two elements
- d) The elements of triads have same valency electrons.

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**Q5. Assertion (A):**

A chemical reaction becomes faster at a higher temperature.

**Reason (R):**

At higher temperature, molecular motion becomes more rapid.

- a) Both (A) and (R) are true, but (R) is not the correct explanation of (A)
- b) (A) is true, but (R) is false
- c) Both (A) and (R) are true and (R) is the correct explanation of (A)
- d) (A) is false, but (R) is true

**Q6.** On moving horizontally across a period, the number of electrons in the outermost shell increases from ..... to .....

- a) 2, 18
- b) 1, 8
- c) 2, 8
- d) 1, 18

**Q7.** Which element forms the highest number of compounds in the periodic table?

- a) Oxygen
- b) Silicon
- c) Carbon
- d) Sulphur

**Q8.** As compared to covalent compounds, electrovalent compounds, generally have

- a) low melting point and high boiling point
  - b) high melting point and low boiling point
  - c) low melting point and low boiling point
  - d) high melting point and high boiling point
- 

**Q9.** What is the correct increasing order of electronegativity for the most common oxidation states among the following elements ?

- a) C N C F
  - b) C N F O
  - c) F O N C
  - d) C N O F
- 

**Q10.** A binary liquid solution is prepared by mixing n-heptane and ethanol. Which of the following statements is incorrect regarding the behaviour of the solution?

- The solution is non-ideal, showing -ve deviation from Raoult's Law.
- The solution is non-ideal, showing +ve deviation from Raoult's Law.
- n-heptane shows +ve deviation while ethanol shows -ve deviation from Raoult's Law.
- The solution formed is an ideal solution.

- a) 1, 2 and 4
  - b) 1, 2 and 3
  - c) 1, 3 and 4
  - d) 2, 3 and 4
- 

**Q11.** Which element is more electronegative among halogens?

- a) Cl
- b) F
- c) Br
- d) I

**Q12.** Which of the following elements A, B, C, D and E with atomic number 3, 11, 15, 18 and 19 respectively belong to the same group?

- a) B, C, D
- b) A, D, E
- c) A, B, C
- d) A, B, E

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**Q13.** The rusting of iron nail

- a) increases its weight
- b) does not affect weight but iron is oxidised
- c) decreases its weight
- d) does not affect weight but iron is reduced

**Q14.** Which of the following properties generally decrease along a period?

- a) Non-metallic character
- b) Metallic character
- c) Atomic size
- d) Both (a) and (c)

**Q15.** Match List-I (Oxidation number) with List-II (the element) and select the correct answer using the code given below the lists

List-I (Oxidation number)	List-II (The element)
A. 2	1. Oxidation number of Mn in $\text{MnO}_2$

B. 3	2. Oxidation number of S in $\text{H}_2\text{SO}_4$
C. 4	3. Oxidation number of Ca in $\text{CaO}$
D. 6	4. Oxidation number of Al in $\text{NaAlH}_4$

**Code:** A B C D

a) 4 3 1 2

b) 3 4 2 1

c) 3 4 1 2

d) 4 3 2 1

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**Answers to the above questions :**

**Q1. Answer: (a)**

All these are noble gases with completely filled outermost shell.

**Q2. Answer: (a)**

**Q3. Answer: (d)**

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**Q4. Answer: (a)**

**Q5. Answer: (c)**

**Q6. Answer: (b)**

**Q7. Answer: (c)**

**Q8. Answer: (d)**

**Q9. Answer: (d)**

**Q10. Answer: (c)**

For this solution intermolecular interactions between n-heptane and ethanol are weaker than n-heptane -n-heptane & ethanol-ethanol interactions hence the solution of n-heptane and ethanol is non-ideal and shows positive deviation from Raoult's law.

**Q11. Answer: (b)**

**Q12. Answer: (d)**

A( $Z = 3$ );

B( $Z = 11$ ) and

E( $Z = 19$ ) are all alkali metals.

**Q13. Answer: (a)**

**Q14. Answer: (d)**

As atomic size decreases along a period valence electrons become more firmly held with the nucleus.

Thus more amount of energy is required to remove valence electrons which reduces the metallic character

**Q15. Answer: (c)**

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