

ATOMIC STRUCTURE, MOLECULES AND NUCLEAR CHEMISTRY BASED GENERAL SCIENCE MCQ PRACTICE QUESTIONS AND ANSWERS PDF WITH EXPLANATION

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Created By [Careericons](#) Team

Q1. Match List-I with List-II and select the correct answer using the code given below the lists.

List-I (Characteristic)	List-II (Particle)
A. Zero mass	1. Positron
B. Fractional charge	2. Neutrino
C. Fractional spin	3. Quark
D. Integral spin	4. Photon

a) 3 2 4 1

b) 2 3 4 1

c) 2 3 1 4

d) 3 2 1 4

Q2. The law of definite proportions was given by –

a) Humphry Davy

b) Proust

c) John Dalton

d) Michael Faraday

Q3. In an atom valence electron are present in

a) next to outermost orbit

b) first orbit

c) outermost orbit

d) any one of its orbit

Q4. The nucleus of an atom contains

- a) electrons
- b) protons and neutrons
- c) protons
- d) neutrons

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Q5. Isotopes are those atoms of the same element which have

- a) Atomic mass is different but atomic number is same
 - b) Atomic number and atoms mass both are same
 - c) Atomic mass is same but atomic number is different
 - d) None of the above
-

Q6. Which of the following are the constituents of matter?

- Atoms
 - Molecules
 - Solid
- a) B and C only
 - b) A and C only
 - c) A and B only
 - d) A, B and C
-

Q7. Molecular mass is defined as the

- a) mass of one atom compared with the mass of one atom of hydrogen

- b) mass of one molecule of any substance compared with the mass of one atom of C-12
 - c) mass of one atom compared with the mass of one molecule
 - d) None of these
-

Q8. The nucleus of a singly ionized carbon atom contains

- a) 6 protons, 6 neutrons and 6 electrons
 - b) 5 protons and 6 neutrons
 - c) 6 protons and 6 neutrons
 - d) 12 protons, 6 neutrons and 6 electrons
-

Q9. The following questions consist of two statements, one labelled as the Assertion (A) and the other as 'Reason (R), You are to examine these two statements carefully and select the answers to these items using the codes given below:

Assertion (A) :

Atomic weights of most of the elements are not whole numbers.

Reason (R) :

Atoms of most of the elements contain mixture of isotopes having different atomic weights.

- a) A is true but R is false
 - b) Both A and R are individually true but R is NOT the correct explanation of A
 - c) Both A and R are individually true and R is the correct explanation of A
 - d) A is false but R is true
-

Q10. Which would be the electrical charge on a sulphur atom containing 18 electrons ?

- a) 1-
 - b) 0
 - c) 2-
 - d) 2+
-

Q11. The atomic weights are expressed in terms of atomic mass unit. Which one of the following is used as a standard?

- a) $^{12}\text{C}_6$
 - b) $^{16}\text{O}_8$
 - c) $^1\text{H}_1$
 - d) $^{35}\text{Cl}_{17}$
-

Q12. Which of the following pairs are isotopes?

- a) Ice and steam
- b) Nitric oxide and nitrogen dioxide
- c) Oxygen and ozone
- d) Hydrogen and deuterium

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Q13. Which of the following pairs is/are correctly matched?

- Isotopes : Atoms with same atomic number but different atomic mass
- Isobars : Atoms with same number of neutrons but different atomic number
- Isotones : Atoms with same mass number but different atomic number

Select the correct answer using the code given below : Code :

- a) 1 and 2 only
 - b) 1 only
 - c) 1, 2 and 3
 - d) 2 only
-

Q14. The cathode ray experiment was done for the first time by

- a) John Dalton

- b) Goldstein
- c) J.J. Thomson
- d) Rutherford

Q15. Which one of the following is heavy water used in nuclear reactor?

- a) Water at 4°C but having molecular weight 19 u
- b) Water having molecular weight 20 u
- c) Water having molecular weight 18 u
- d) Water below the ice in a frozen sea

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Answers to the above questions :

Q1. Answer: (c)

Q2. Answer: (b)

Q3. Answer: (c)

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Q4. Answer: (b)

The nucleus contains protons and neutrons whereas electrons revolves around the nucleus in circular orbits.

Q5. Answer: (a)

Q6. Answer: (c)

Q7. Answer: (b)

Q8. Answer: (c)

Q9. Answer: (c)

Q10. Answer: (c)

S(16) = 2, 8, 6

Hence S(18) need two or more electron to complete its octet.

i.e $S + 2e^- \rightarrow S^{2-}$

Q11. Answer: (a)

$^{12}_6C$ used as a standard in the expression of atomic weights in term of amu.

Q12. Answer: (d)

Q13. Answer: (b)

Q14. Answer: (c)

Q15. Answer: (b)

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