

GENERAL CONCEPTS OF CHEMISTRY BASED GENERAL SCIENCE MCQ PRACTICE QUESTIONS AND ANSWERS PDF WITH EXPLANATION

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Created By [Careericons](#) Team

Q1. $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$ Above reaction is -

- a) combination reaction
 - b) precipitation reaction
 - c) double displacement reaction
 - d) Both (b) and (c)
-

Q2. Which is not a balanced equation ?

- a) $\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$
 - b) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
 - c) $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$
 - d) $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$
-

Q3. Which of the following does not corrode when exposed to the atmosphere?

- a) Gold
 - b) Iron
 - c) Copper
 - d) Silver
-

Q4. When electrons are added the resulting ion is called

- a) neutral radical
- b) basic radical
- c) acidic radicals

d) None of these

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Q5. The empirical formula and molecular mass of a compound are CH_2O and 180 g respectively. What will be the molecular formula of the compound ?

a) $\text{C}_6\text{H}_{12}\text{O}_6$

b) $\text{C}_9\text{H}_{18}\text{O}_9$

c) CH_2O

d) $\text{C}_2\text{H}_4\text{O}_2$

Q6. The observation that does not indicate a chemical reaction is

a) evolution of a gas

b) change of temperature

c) change of colour

d) None of these

Q7. Which of the following statements is true about chemical equation ?

a) Mass as well as atoms are conserved

b) Mass is conserved

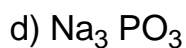
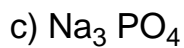
c) Atoms are conserved

d) Molecules are conserved

Q8. The formula of sodium pyrophosphate is

a) $\text{Na}_4\text{P}_2\text{O}_7$

b) $\text{Na}_2\text{P}_2\text{O}_7$



Q9. The branch of chemistry which deals with study of physical properties and conditions is

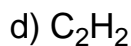
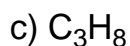
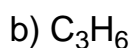
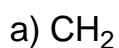
a) nuclear chemistry

b) physical chemistry

c) analytical chemistry

d) pharmaceutical chemistry

Q10. The empirical formula of a compound is CH_2 . One mole of this compound has a mass of 42 grams. Its molecular formula is :



Q11. Rusting of an iron is an example of

a) oxidation

b) reduction

c) ionization

d) dissociation

Q12. Which one of the following is not an exothermic reaction ?

a) Burning of a candle

b) Respiration

c) Slaking of lime

d) Dipping of iron block in water

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Q13. A molal solution is one that contains 1 mole of a solute in

- a) one litre of the solution
- b) one litre of the solvent
- c) 1000 g of the solvent
- d) 22.4 litres of the solution

Q14. The reactions in which two compounds exchange their radicals to form two new compounds are called

- a) double displacement reaction
- b) displacement reaction
- c) decomposition reaction
- d) isomerisation reaction

Q15. Which of the following involves combination of two elements?

- a) $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$
- b) $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
- c) $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
- d) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$

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Answers to the above questions :

Q1. Answer: (d)

This reaction is double displacement and precipitation as well because insoluble silver chloride gets precipitated.

Q2. Answer: (a)

$2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$ (balanced chemical eqn.)

Q3. Answer: (a)

Gold is a noble metal.

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Q4. Answer: (c)

$\text{Cl} + \text{e}^- \text{Cl}^-$ (acidic radicals)

Q5. Answer: (a)

$n = \frac{180}{30}$ Molecular mass of CH_2O = 180/30

n=6

Q6. Answer: (d)

A chemical reaction is generally accompanied by a change of temperature, colour or evolution of a gas.

Q7. Answer: (a)

Q8. Answer: (a)

Q9. Answer: (b)

Q10. Answer: (b)

Empirical formula of compound = CH_2

Molecular mass of the compound = 42

? $n = 42/14 = 3$

? Hence molecular formula = C_3H_6

Q11. Answer: (a)

Rusting results in formation of iron oxide.

Q12. Answer: (d)

No reaction occurs in this case.

Q13. Answer: (c)

Molal solution contains 1 mole of solute in 1000 g solvent.

Q14. Answer: (a)

Q15. Answer: (c)

Both Na and O₂ are elements.

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