

# GENERAL CONCEPTS OF CHEMISTRY BASED GENERAL SCIENCE MCQ PRACTICE QUESTIONS AND ANSWERS PDF WITH EXPLANATION

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**Q1.** Chemical formula of Aluminium sulphate is

- a)  $\text{Al}_3(\text{SO}_4)_2$
  - b)  $\text{Al}_2(\text{SO}_4)_3$
  - c)  $\text{AlSO}_4$
  - d) None of these
- 

**Q2.** The molecular formula  $\text{P}_2\text{O}_5$  means that

- a) there are twice as many P atoms in the molecule as there are O atoms
  - b) a molecule contains 2 atoms of P and 5 atoms of O
  - c) the ratio of the mass of P to the mass of O in the molecule is 2:5
  - d) the ratio of the mass of P to the mass of O in the molecule is 5 : 2
- 

**Q3.** Which one is a bivalent ion?

- a) sulphide
  - b) sodium
  - c) calcium
  - d) Both (b) and (c)
- 

**Q4.** In the reaction :  $\text{PbO} + \text{C} \rightarrow \text{Pb} + \text{CO}$

- a) C is reductant
- b) PbO is oxidised
- c) PbO is oxidant

d) Both (a) and (c)

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**Q5.** The branch of chemistry which deals with study of the methods of detection and determination of elements and compounds is

- a) Analytical chemistry
- b) Physical chemistry
- c) Nuclear chemistry
- d) Bio chemistry

**Q6.** Which of the following statement is not correct ? To prevent food from getting rancid, we

- a) flush bags of chips with an inert gas
- b) store the food in air-tight containers
- c) add antioxidants
- d) increase the temperature of the food

**Q7.** Which one is a combination reaction ?

- a)  $\text{Ca(OH)}_2 + \text{Na}_2\text{CO}_3 \rightarrow 2\text{NaOH} + \text{CaCO}_3$  ?
- b) Formation of a mixture of carbon monoxide and hydrogen. When steam is passed over red hot iron.
- c) Reaction of water with sodium metal to form sodium hydroxide and hydrogen.
- d) Preparation of stannic chloride (Tin (iv) chloride) by passing chlorine gas into molten tin (Sn).

**Q8.** Which of the following reaction is endothermic?

- a)  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$   
b)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$   
c)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
d)  $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
- 

**Q9.** Valency of an atom or radicals is

- a) its combining capacity  
b) ionisation energy  
c) electron affinity of atom  
d) size of atom
- 

**Q10.** An organic compound containing C, H and O gave on analysis C – 40% and H – 6.66%. Its empirical formula would be

- a)  $\text{CH}_2\text{O}$   
b)  $\text{C}_3\text{H}_6\text{O}$   
c)  $\text{CHO}$   
d)  $\text{CH}_4\text{O}$
- 

**Q11.** Which one is a decomposition reaction ?

- a)  $2\text{H}_2\text{O} \xrightarrow{\text{Electrolysis}} 2\text{H}_2 + \text{O}_2$   
b)  $2\text{HgO} \xrightarrow{\text{heat}} 2\text{Hg} + \text{O}_2$   
c)  $\text{CaCO}_3 \xrightarrow{\text{heat}} \text{CaO} + \text{CO}_2$   
d) All the above reactions are decomposition reaction
- 

**Q12.**  $\text{Na}_2\text{S}_2\text{O}_3$  represent the compound

- a) sodium thiosulphate  
b) sodium sulphate  
c) sodium sulphite

d) None of these

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**Q13.** The chromate and dichromate ions are respectively

- a)  $\text{CrO}_4^-$  and  $\text{CrO}_5^-$
- b)  $\text{CrO}_4^{2-}$  and  $\text{Cr}_2\text{O}_7^{2-}$
- c)  $\text{Cr}_2\text{O}_7^{2-}$  and  $\text{CrO}_4^-$
- d)  $\text{CrO}_4^{2-}$  and  $\text{Cr}_2\text{O}_5^{2-}$

**Q14.** A redox reaction is one in which -

- a) an acid is neutralised by the base
- b) both the substance are reduced
- c) both the substance are oxidised
- d) one substance is oxidised while the other is reduced

**Q15.** Slow eating away of iron articles in the presence of moist air is called

- a) rusting
- b) galvanisation
- c) crystallisation
- d) neutralisation

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**Answers to the above questions :**

**Q1. Answer: (b)**

**Q2. Answer: (b)**

**Q3. Answer: (d)**

Ca<sup>++</sup> and S<sup>--</sup>

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**Q4. Answer: (d)**

In this carbon reduces PbO to Pb so carbon (C) is reducing agent (reductant) PbO acts as oxidising agent (oxidant) as it oxidises C to CO.

**Q5. Answer: (a)**

**Q6. Answer: (d)**

Increase the temperature of the food.

**Q7. Answer: (d)**

$\text{Sn} + 2\text{Cl}_2 \rightarrow \text{SnCl}_4$ .

**Q8. Answer: (c)**

Heat is required to decompose calcium carbonate. Thus this reaction proceeds with absorption of heat therefore it is endothermic reaction.

**Q9. Answer: (a)**

**Q10. Answer: (a)**

Element	%	At.Wt.	Rel.Number	Ratio
C	40	12	$\frac{40}{12} = 3.33$	$\frac{3.33}{3.33} = 1$
H	6.66	1	$\frac{6.66}{1} = 6.66$	$\frac{6.66}{3.33} = 2$
O	53.34	16	$\frac{53.34}{16} = 3.33$	$\frac{3.33}{3.33} = 1$

(% of O in organic compound  
=  $100 - (40 + 6.66) = 53.34\%$  )

Empirical formula of organic compound =  $CH_2O$ .

**Q11. Answer: (d)**

In all these reactions the reactants decompose to form simpler products.

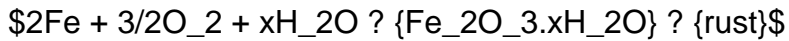
**Q12. Answer: (a)**

**Q13. Answer: (b)**

**Q14. Answer: (d)**

**Q15. Answer: (a)**

Rusting is a process in which iron gets converted into hydrated iron oxide in presence of moisture.



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